

## Review: Naming and Writing Formulas for Ionic Compounds, Covalent Molecules, and Acids

Name \_\_\_\_\_ Date \_\_\_\_\_ Period \_\_\_\_\_

**Writing Formulas:** Combine each cation with each anion to write the correct formula for the ionic compound.

	Cl <sup>-</sup>	CO <sub>3</sub> <sup>2-</sup>	OH <sup>-</sup>	SO <sub>4</sub> <sup>2-</sup>	PO <sub>4</sub> <sup>3-</sup>	NO <sub>3</sub> <sup>-</sup>
Na <sup>+</sup>						
NH <sub>4</sub> <sup>+</sup>						
K <sup>+</sup>						
Ca <sup>2+</sup>						
Mg <sup>2+</sup>						
Zn <sup>2+</sup>						
Fe <sup>3+</sup>						
Al <sup>3+</sup>						
Co <sup>3+</sup>						
Fe <sup>2+</sup>						
H <sup>+</sup>						

**Naming Ionic Compounds**

CaCO <sub>3</sub>	
KCl	
*FeSO <sub>4</sub>	
LiBr	
MgCl <sub>2</sub>	
*FeCl <sub>3</sub>	
Zn <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub>	
NH <sub>4</sub> NO <sub>3</sub>	
Al(OH) <sub>3</sub>	
*CuC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	
*PbSO <sub>3</sub>	
NaClO <sub>3</sub>	
CaC <sub>2</sub> O <sub>4</sub>	
*Fe <sub>2</sub> O <sub>3</sub>	
(NH <sub>4</sub> ) <sub>3</sub> PO <sub>4</sub>	
NaHSO <sub>4</sub>	
*Hg <sub>2</sub> Cl <sub>2</sub>	
Mg(NO <sub>2</sub> ) <sub>2</sub>	
*CuSO <sub>4</sub>	
NaHCO <sub>3</sub>	
*NiBr <sub>3</sub>	
Be(NO <sub>3</sub> ) <sub>2</sub>	
ZnSO <sub>4</sub>	
*AuCl <sub>3</sub>	
KMnO <sub>4</sub>	

\*Requires use of Stock System for naming

## Naming and Writing Formulas for Molecular Compounds

1.	CO <sub>2</sub>	
2.	CO	
3.	SO <sub>2</sub>	
4.	SO <sub>3</sub>	
5.	N <sub>2</sub> O	
6.	NO	
7.	N <sub>2</sub> O <sub>3</sub>	
8.	NO <sub>2</sub>	
9.	N <sub>2</sub> O <sub>4</sub>	
10.	N <sub>2</sub> O <sub>5</sub>	
11.	PCl <sub>3</sub>	
12.	PCl <sub>5</sub>	
13.	NH <sub>3</sub>	
14.	SCl <sub>6</sub>	
15.	P <sub>2</sub> O <sub>5</sub>	
16.	CF <sub>4</sub>	
17.	SiO <sub>2</sub>	
18.	CS <sub>2</sub>	
19.	OF <sub>2</sub>	
20.	PBr <sub>3</sub>	
21.		sulfur hexafluoride
22.		nitrogen triiodide
23.		water
24.		ammonia
25.		xenon pentafluoride
26.		carbon tetrachloride
27.		antimony pentachloride
28.		bromine pentafluoride
29.		chlorine trifluoride
30.		silicon tetrachloride

## Naming Acids

1.	HNO <sub>3</sub>	
2.	HCl	
3.	H <sub>2</sub> SO <sub>4</sub>	
4.	H <sub>2</sub> SO <sub>3</sub>	
5.	HC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	
6.	HBr	
7.	HNO <sub>2</sub>	
8.	H <sub>3</sub> PO <sub>4</sub>	
9.	H <sub>2</sub> S	
10.	H <sub>2</sub> CO <sub>3</sub>	
11.	H <sub>2</sub> Se	
12.	HMnO <sub>4</sub>	

**Writing Formulas:** Write the formula for the following acids.

13.	Sulfuric acid	
14.	Nitric acid	
15.	Hydrochloric acid	
16.	Acetic acid	
17.	Hydrofluoric acid	
18.	Phosphorous acid	
19.	Carbonic acid	
20.	Nitrous acid	
21.	Phosphoric acid	
22.	Hydrosulfuric acid	
23.	Perchloric acid	
24.	Oxalic acid	