

Covalent Bonding

Section 5 Electronegativity and Polarity

Main Idea

Details

Scan Section 5 of your text. Use the checklist below as a guide.

- Read all section titles.
- Read all boldfaced words.
- Read all tables and charts.
- Look at all pictures and read the captions.
- Think about what you already know about the strengths and distribution of charge in covalent bonds.

Write three facts you discovered about electronegativity.

1. _____
2. _____
3. _____

Review Vocabulary

Use your brain, notes, and/or textbook to define the following terms.

covalent bond

molecule

nonpolar covalent bond

New Vocabulary

Use your text to define the following term.

polar covalent bond

Section 5 Electronegativity and Polarity (continued)

Main Idea

Electronegativity and Bond Character

Use with pages 265–265.

Polar Covalent Bonds

Use with pages 267–268.

Details

Sequence *the following elements from the least electronegative to the most electronegative. Use Figure 20 for reference.*

_____ Au

_____ Y

_____ Ba

_____ P

_____ H

_____ Te

_____ O

_____ I

_____ Co

Draw *the Lewis structure for each of the molecular compounds listed below. Analyze the symmetry of the structure to determine whether or not the compound is polar covalent or nonpolar covalent.*

N₂ _____

CO₂ _____

CH₃Cl _____